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Distillation is a process in which a liquid mixture of volatile components is separated by imparting energy to it in consideration with the boiling points of the components so that selective vaporization takes place. This process can also be used in reverse to selectively condense the vapour mixture.

What Is Distillation? - Chemical Engineering World

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Z. Lei, in Reference Module in Chemistry, Molecular Sciences and Chemical Engineering, 2017. Introduction. Distillation is the most commonly used method for the separation of homogeneous liquid mixtures and is based on differences in the boiling points or relative volatility of the constituent components. 1,2 However, where the relative volatility is close to unity (e.g., for the separation of azeotropic mixtures or close boiling components), a third component (i.e., entrainer, solvent, or ...

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Extractive Distillation An alternative to recover ethanol is to use extractive distillation. The solvent used is Propylene Glycol. Recall also that ethanol forms a minimum-boiling azeotrope with ...

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Distillation refers to the selective boiling and subsequent condensation of a component in a liquid mixture. It is a separation technique that can be used to either increase the concentration of a particular component in the mixture or to obtain (almost) pure components from the mixture.

Distillation - Definition, Detailed Process, Types, Uses

Distillation is the process of separating the components or substances from a liquid mixture by using selective boiling and condensation. Distillation may result in essentially complete separation (nearly pure components), or it may be a partial separation that increases the concentration of selected components in the mixture.

Distillation - Wikipedia

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